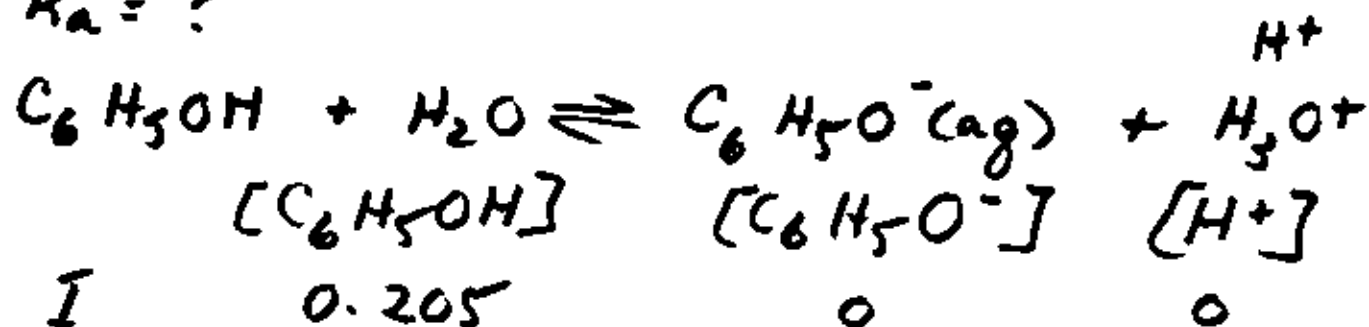


$$3/ [C_6H_5OH] = 0.205$$

$$[H^+] = 2.340 \times 10^{-6}$$

$$K_a = ?$$



C	-2.340×10^{-6}	2.340×10^{-6}	2.34×10^{-6}
---	-------------------------	------------------------	-----------------------

E	$\frac{0.204997}{0.205}$	2.340×10^{-6}	2.34×10^{-6}
---	--------------------------	------------------------	-----------------------

$$K_a = \frac{[2.340 \times 10^{-6}]^2}{0.205} = 2.67 \times 10^{-11}$$